

(Multiple Choice Questions)

This section contains **90 multiple choice questions**. Each question has 4 choices (1), (2), (3) and (4) for its answer, out of which **ONLY ONE** is correct. (+4, -1)

- The formation of the oxide ion, O^{2-} (g), from oxygen atom requires first an exothermic and then an endothermic step as shown below:
 $O(g) + e^- \longrightarrow O^-(g); \Delta_{eg}H = -141 \text{ kJ/mol}$
 $O^-(g) + e^- \longrightarrow O^{2-}(g); \Delta_{eg}H = +780 \text{ kJ/mol}$
Thus process of formation of O^{2-} in gas phase is unfavourable even though O^{2-} is isoelectronic with neon. It is due to the fact that
(1) Oxygen is more electronegative.
(2) Addition of electron in oxygen results in larger size of the ion.
(3) Electron repulsion outweighs the stability gained by achieving noble gas configuration.
(4) O^- ion has comparatively smaller size than oxygen atom.
- From B_2H_6 , all the following can be prepared except –
(1) B_2O_3 (2) H_3BO_3 (3) $B_2(CH_3)_6$ (4) $NaBH_4$
- Name the metal M which is extracted on the basis of following reactions
 $4M + 8CN^- + 2H_2O + O_2 \rightarrow 4[M(CN)_2]^- + 4OH^-$
 $2[M(CN)_2]^- + Zn \rightarrow [Zn(CN)_4]^{2-} + 2M$
(1) Ni (2) Ag (3) Cu (4) Hg
- Poling process is used for
(1) the removal of Cu_2O from Cu (2) the removal of Al_2O_3 from Al
(3) the removal of Fe_2O_3 from Fe (4) All of these
- Which complex is likely to show optical activity
(1) Trans $-[Co(NH_3)_4Cl_2]^+$ (2) $[Cr(H_2O)_6]^{3+}$
(3) Cis $-[Co(NH_3)_2(en)_2]^{3+}$ (4) Trans $-[Co(NH_3)_2(en)_2]^{3+}$
- On adding KI to a metal salt solution, no precipitate was observed but the salt solution gives yellow precipitate with K_2CrO_4 in the presence CH_3COOH . Then the salt is
(1) $Sr(NO_3)_2$ (2) $Pb(CH_3COO)_2$ (3) $AgNO_3$ (4) $BaCl_2$
- Cinnabar is
(1) CuS (2) Ag_2S (3) ZnS (4) HgS
- Smog is essentially caused by the presence of
(1) O_3 and N_2 (2) O_2 and N_2
(3) Oxides of sulphur and N_2 (4) O_2 and O_3

9. Which of the following is a lanthanoids
 (1) Ta (2) Rh (3) Th (4) Lu
10. Which of the following will not be oxidised by O_3 ?
 (1) KI (2) $FeSO_4$ (3) $KMnO_4$ (4) K_2MnO_4
11. Which of the following compound is not an antacid?
 (1) Aluminium hydroxide (2) Cimetidine
 (3) Phenelzine (4) Ranitidine
12. Which one of the following pairs of isomers and types of isomerism are correctly matched?
 a. $[Co(NH_3)_5(NO_2)]Cl_2$ and $[Co(NH_3)_5(ONO)]Cl_2$ Linkage
 b. $[Cu(NH_3)_4]PtCl_4$ and $[Pt(NH_3)_4][CuCl_4]$ Co – ordination
 c. $[Pt(NH_3)_4Cl_2]Br_2$ and $[Pt(NH_3)_4Br_2]Cl_2$ Ionization
 Select the correct answer using the codes given below:
 (1) b and c (2) a,b and c (3) a and c (4) a and b
13. The gas leaked from a storage tank of the union carbide plant in Bhopal gas tragedy was
 (1) Ammonia (2) Phosgene (3) Methyl isocyanate (4) Methyl amine
14. Which is low spin complex?
 (1) $Fe(CN)_6^{3-}$ (2) $Co(NO_2)_6^{3-}$ (3) $Mn(CN)_6^{3-}$ (4) All
15. All fluorine atoms are in same plane in
 (1) CHF_3 (2) ClF_3 (3) $XeOF_4$ (4) All of these
16. Which of the following oxides of chromium is amphoteric in nature
 (1) CrO (2) Cr_2O_3 (3) CrO_3 (4) CrO_5
17. Pure Cl_2 is prepared on heating
 (1) $NaCl$ (2) $PtCl_4$ (3) $CuCl_2$ (4) All of these
18. The set of d-orbitals which do not contain any d-orbital which is involved in hybridization of central atom in ICl_4^-
 (1) $d_{z^2}, d_{x^2-y^2}$ (2) $d_{x^2-y^2}, d_{xy}, d_{yz}, d_{zx}$ (3) $d_{z^2}, d_{xy}, d_{yz}, d_{zx}$ (4) d_{xy}, d_{zx}, d_{yz}
19. Leaching is a process of
 (1) Reduction (2) Concentration (3) Refining (4) Oxidation
20. Out of N, O, Ne, Na, Na^+ , select the species which have maximum and minimum ionization energy
 (1) Na^+, Na (2) NO, Na (3) Na^+, O (4) Ne, N
21. The correct ionic radii order is
 (1) $N^{3-} > O^{2-} > F^- > Na^+$ (2) $N^{3-} > Na^+ > O^{2-} > F^-$
 (3) $Na^+ > O^{2-} > N^{3-} > F^-$ (4) $O^{2-} > F^- > Na^+ > N^{3-}$
22. H_3BO_3 is
 (1) Mono basic and weak Lewis acid (2) Mono basic and weak Bronsted acid
 (3) Mono basic acid and strong Lewis acid (4) Tri basic acid and weak Bronsted acid

23. Which metal can be purified by distillation?
 (1) Cu (2) Ag (3) Fe (4) Hg
24. Chemical A is used for water softening to remove temporary hardness. A reacts with Na_2CO_3 to generate caustic soda. When CO_2 is bubbled through A, it turns cloudy. What is the chemical formula of A :
 (1) CaCO_3 (2) CaO (3) Ca(OH)_2 (4) $\text{Ca(HCO}_3)_2$
25. A metal M readily forms water soluble sulphate MSO_4 , water insoluble hydroxide M(OH)_2 and oxide MO which becomes inert on heating. The hydroxide is soluble in NaOH . The M is
 (1) Be (2) Mg (3) Ca (4) Sr
26. Which of the following facts regarding boron and silicon is not true?
 (1) Boron is used to make boron steel or boron carbide control rods for nuclear reactor.
 (2) Boron and silicon form halides which are not hydrolysed.
 (3) Boron and silicon react with magnesium to form magnesium boride and magnesium silicide which are decomposed by acids to give volatile borane and silane, respectively.
 (4) Both boron and silicon react with alkali to form borates and silicates containing BO_3^{3-} and SiO_4^{4-} tetrahedral units, respectively.
27. In electrolysis of NaCl when Pt electrode is taken then H_2 is liberated at cathode while with Hg cathode it forms sodium amalgam
 (1) Hg is more inert than Pt
 (2) More voltage is required to reduce H^+ at Hg than at Pt
 (3) Na is dissolved in Hg while it does not dissolve in Pt
 (4) Conc. of H^+ ions is larger when Pt electrode is taken
28. Stable oxide is obtained by heating the carbonate of the element
 (1) Li (2) K (3) Na (4) Rb
29. Nitric oxide is
 (1) Acidic towards litmus (2) Basic towards litmus
 (3) Neutral towards litmus (4) Amphoteric
30. Which of the following is incorrect?
 (1) Mg burns in air releasing dazzling light rich in UV rays.
 (2) $\text{CaCl}_2 \cdot 6\text{H}_2\text{O}$ when mixed with ice gives freezing mixture.
 (3) Mg cannot form complexes
 (4) Be can form complexes due to its very small size.
31. Which one has the highest bond energy?
 (1) O—O (2) S—S (3) Se—Se (4) Te—Te
32. Which of the following statement is not correct
 (1) LiOH is amphoteric in nature
 (2) LiCl is soluble in pyridine
 (3) Li_3N is stable while Na_3N doesn't exist even at room temperature
 (4) BeO is amphoteric in nature
33. Potassium manganate (K_2MnO_4) is formed when:
 (1) Cl_2 is passed into an aqueous KMnO_4 solution
 (2) MnO_2 is fused with KOH
 (3) Formaldehyde reacts with KMnO_4 in presence of strong alkali
 (4) KMnO_4 reacts with concentrated H_2SO_4

34. Which of the following is covalent carbide?
 (1) CaC_2 (2) Al_4C_3 (3) SiC (4) Be_2C
35. (White ppt) $\text{D} \xrightarrow{\text{Na}_2\text{CO}_3} \text{A} \xrightarrow[\text{(in acetic acid)}]{\text{K}_2\text{CrO}_4} \text{B}$ (Yellow ppt)
 \downarrow
 dil. H_2SO_4
 C (White ppt)
 If A is the metallic salt, then the white ppt. of D must be of
 (1) Magnesium oxide (2) Red lead
 (3) Barium carbonate (4) Calcium carbonate
36. The more commonly used baking powder contains about 30% NaHCO_3 , 20% $\text{NaAl}(\text{SO}_4)_2$, 10% $\text{Ca}(\text{H}_2\text{PO}_4)_2$ and 40% starch. Which of the following statements are correct?
 (1) $\text{Ca}(\text{H}_2\text{PO}_4)_2$ is acidic and when moistened it reacts with NaHCO_3 evolving CO_2 gas
 (2) $\text{NaAl}(\text{SO}_4)_2$ slows down the decomposition reaction of NaHCO_3 so that CO_2 is evolved more slowly.
 (3) $\text{NaAl}(\text{SO}_4)_2$ is acidic and when moistened it reacts with NaHCO_3 evolving CO_2
 (4) Both (A) and (B)
37. $\text{Mg}_2\text{C}_3 + \text{H}_2\text{O} \longrightarrow \text{X}$ (organic compound). Compound X is
 (1) C_2H_2 (2) CH_4 (3) Propyne (4) Ethane
38. The final product obtained when boric acid is heated to red heat is –
 (1) Metaboric acid (2) Tetraboric acid (3) Boron oxide (4) Pyroboric acid
39. Select the correct order of basic strength?
 (1) $\text{KOH} > \text{Mg}(\text{OH})_2 > \text{Al}(\text{OH})_3$ (2) $\text{Al}(\text{OH})_3 > \text{Mg}(\text{OH})_2 > \text{KOH}$
 (3) $\text{Mg}(\text{OH})_2 > \text{KOH} > \text{Al}(\text{OH})_3$ (4) $\text{KOH} > \text{Al}(\text{OH})_3 > \text{Mg}(\text{OH})_2$
40. Silicones have the general formula –
 (1) SiO_4^{4-} (2) $\text{Si}_2\text{O}_7^{6-}$ (3) $(\text{R}_2\text{SiO})_n$ (4) NaBH_4
41. Which of the following statement is false?
 (1) Optical isomerism is observed in $[\text{M}_{abcd}]^{n\pm}$ tetrahedral complexes.
 (2) Geometrical isomerism does not exist while optical isomerism exists in the complex $[\text{Fe}(\text{C}_2\text{O}_4)_3]^{3-}$
 (3) Both *cis* and *trans* forms are optically inactive in $[\text{PtCl}_2(\text{NH}_3)_4]^{2+}$ complex ion.
 (4) $[\text{Pt}(\text{en})_2]^{2+}$ shows geometrical isomerism as well as optical.
42. The species present in solution when CO_2 is dissolved in water are –
 (1) $\text{CO}_2, \text{H}_2\text{CO}_3, \text{HCO}_3^-, \text{CO}_3^{2-}$ (2) $\text{H}_2\text{CO}_3, \text{CO}_3^{2-}$
 (3) $\text{HCO}_3^-, \text{CO}_3^{2-}$ (4) $\text{CO}_2, \text{H}_2\text{CO}_3$
43. Element X, Y and Z have atomic nos. 19, 37 and 55 respectively which of the following statement is true
 (1) Their ionisation potential would increase with the increasing atomic number.
 (2) Y would have an ionisation potential in between those of X and Z
 (3) Z would have the highest I.P.
 (4) Y would have the highest I.P.

44. What would happen when a small quantity of H_2O_2 is added to a solution of FeSO_4 ?
 (1) Colour of FeSO_4 disappears. (2) H_2 is evolved.
 (3) An electron is added to Fe^{2+} . (4) An electron is lost by Fe^{2+} .
45. The dissolution of $\text{Al}(\text{OH})_3$ by a solution of NaOH results in the formation of
 (1) $[\text{Al}(\text{H}_2\text{O})_4(\text{OH})]^{2+}$ (2) $[\text{Al}(\text{H}_2\text{O})_2(\text{OH})_4]^-$
 (3) $[\text{Al}(\text{H}_2\text{O})_3(\text{OH})_3]$ (4) $[\text{Al}(\text{H}_2\text{O})_6(\text{OH})_3]$
46. On the addition of mineral acid to an aqueous solution of borax, the compound formed is:
 (1) Borodihydride (2) Orthoboric acid (3) Metaboric acid (4) Pyroboric acid
47. Which of the following statement is correct for diborane?
 (1) Small amines like NH_3 , CH_3NH_2 give unsymmetrical cleavage of diborane.
 (2) Large amines such as $(\text{CH}_3)_3\text{N}$ and pyridine gives symmetrical cleavage of diborane.
 (3) Small as well as large amines both gives symmetrical cleavage of diborane.
 (4) (A) and (B) both
48. Boron compounds behave as Lewis acid because of
 (1) Acidic nature (2) Covalent nature (3) Ionic nature (4) Vacant orbital
49. Consider the following molecules, NO , CO and O_2 select incorrect statement about these
 (1) O_2 is paramagnetic and, CO and NO are diamagnetic.
 (2) NO and O_2 are paramagnetic, CO is diamagnetic
 (3) Bond order is in the order, $\text{O}_2 < \text{NO} < \text{CO}$.
 (4) Number of unpaired electrons are 1, 0, 2 respectively
50. The shape and hybridization of central atom of compound $\text{N}(\text{SiH}_3)_3$
 (1) planar, sp^2 (2) tetrahedral, sp^3 (3) bent, sp^3 (4) trigonal bipyramidal, sp^3d
51. An alkaline earth metal (M) gives a salt with chlorine, which is soluble in water at room temperature. It also forms an insoluble sulphate whose mixture with a sulphide of a transition metal is called 'lithopone' -a white pigment. Metal M is
 (1) Ca (2) Mg (3) Ba (4) Sr
52. $\text{p}\pi - \text{p}\pi$ bonding is not present in
 (1) NO_3^- (2) NO_2^- (3) N_3^- (4) PO_4^{3-}
53. Which halide does not hydrolyse
 (1) SbCl_3 (2) AsCl_3 (3) PCl_3 (4) NF_3
54. Which of the following statements is wrong?
 (1) An acidified solution of $\text{K}_2\text{Cr}_2\text{O}_7$ liberates iodine from iodides
 (2) In acidic solution dichromate ions are converted to chromate ions
 (3) Ammonium dichromate on heating undergoes exothermic decomposition to give Cr_2O_3 & N_2
 (4) Potassium dichromate is used as a titrant for Fe^{2+} ions
55. Roasting is used in the extraction of
 (1) Galena (2) Iron pyrite (3) Copper glance (4) All of these
56. In which reaction does SO_2 acts as oxidizing agent
 (1) Acidified KMnO_4 (2) Acidified $\text{K}_2\text{Cr}_2\text{O}_7$
 (3) Acidified $\text{C}_2\text{H}_5\text{OH}$ (4) H_2S

57. Consider $M(OH)_3$ formed by all the group 13 elements. The correct sequence of acidic strength of hydroxides $[M(OH)_3]$ is
 (1) $B(OH)_3 < Al(OH)_3 > Ga(OH)_3 > In(OH)_3 > Tl(OH)_3$
 (2) $B(OH)_3 > Tl(OH)_3 > Al(OH)_3 > In(OH)_3 > Ga(OH)_3$
 (3) $Al(OH)_3 > Ga(OH)_3 > B(OH)_3 > In(OH)_3 > Tl(OH)_3$
 (4) $B(OH)_3 > Al(OH)_3 > Ga(OH)_3 > In(OH)_3 > Tl(OH)_3$
58. The lanthanoids contraction relates to
 (1) Atomic radii (2) Atomic as well as M^{3+} radii
 (3) Valence electrons (4) Oxidation states
59. Cis-Trans-isomerism is found in square planar complexes of the molecular formula (a and b are monodentate ligands)
 (1) Ma_4 (2) Ma_3b (3) Ma_2b_2 (4) Mab_3
60. Which reaction is possible
 (1) $I_2 + 2NaBr \rightarrow Br_2 + 2NaI$ (2) $I_2 + 2NaCl \rightarrow Cl_2 + 2NaI$
 (3) $Br_2 + 2NaCl \rightarrow Cl_2 + 2NaBr$ (4) $Cl_2 + 2NaBr \rightarrow Br_2 + 2NaCl$
61. Which complex has square planar shape?
 (1) $[Ni(CN)_4]^{2-}$ (2) $[Cu(NH_3)_4]^{2+}$ (3) $[PtCl_4]^{2-}$ (4) All of these
62. Decreasing order of reducing power hydrogen halides is
 (1) $HI > HBr > HCl > HF$ (2) $HF > HI > HBr > HCl$
 (3) $HI > HF > HBr > HCl$ (4) None
63. Consider a P_y orbital of an atom and identify correct statement.
 (1) s-orbital of another atom produce π bond when y is the bond formation axis.
 (2) P_y - orbital of another produce π bond when x is the bond formation axis.
 (3) P_z - orbital of another atom produce π bond when x is the bond formation axis.
 (4) d_{xy} - orbital of another atom produce π bond when x is the bond formation axis.
64. The element which forms oxides in all the oxidation states from +1 to +5, is;
 (1) P (2) Sb (3) N (4) As
65. The geometry, hybridization and magnetic moment of $[Ni(CO)_4]$ are.....andrespectively
 (1) tetrahedral, sp^3 , zero (2) square planar, sp^3 , $\sqrt{3}$
 (3) tetrahedral, sp^3 , $2\sqrt{2}$ (4) square planar, dsp^2 , zero
66. Increasing order of acidic strength of oxy-acids of chlorine is:
 (1) $HClO < HClO_3 < HClO_2 < HClO_4$ (2) $HClO_4 < HClO_2 < HClO < HClO_3$
 (3) $HClO < HClO_2 < HClO_3 < HClO_4$ (4) None of these
67. Which of the following process is used in the extractive metallurgy of Mg?
 (1) Fused salt electrolysis (2) Self - reduction
 (3) Aqueous solution electrolysis (4) Thermite reduction
68. Which transition elements exhibit +8 oxidation state:
 (1) Cu, Zn (2) Ru, Os (3) Ag, Au (4) Cu, Cr
69. Which of the following is not tranquilizer?
 (1) Luminal (2) Seconal (3) Reserpine (4) Piperazine

70. The following equation represents a method of purification of nickel by,
- $$\underset{\text{Im pure}}{\text{Ni}} + 4\text{CO} \xrightarrow{320\text{K}} \text{Ni}(\text{CO})_4 \xrightarrow{420\text{K}} \underset{\text{pure}}{\text{Ni}} + 4\text{CO}$$
- (1) Cupellation (2) Mond's process
(3) Van Arkel method (4) Zone refining
71. $\underset{\text{(Black)}}{\text{X}} \xrightarrow{\text{Dil. H}_2\text{SO}_4} \underset{\text{(Gas)}}{\text{Y}} \xrightarrow{\text{Dil. HNO}_3} \text{Colloidal sulphur}$ Identify X?
- (1) CuS (2) FeS (3) As₂S₃ (4) CdS
72. The colour of $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ is due to
- (1) Transfer of an electron from one Ti to another
(2) Presence of water molecule
(3) Excitation of electrons from d → d
(4) Intramolecular vibration
73. Select the correct order of hydration energy of ions?
- (1) Li⁺ > Na⁺ > K⁺ (2) Li⁺ > K⁺ > Na⁺ (3) Cs⁺ > Rb⁺ > K⁺ (4) Li⁺ > Cs⁺ > K⁺
74. Magnesium on reaction with very dilute HNO₃ gives
- (1) NO (2) N₂O (3) H₂ (4) NO₂
75. The pair of compounds having metals in their highest oxidation state
- (1) MnO₂, FeCl₃ (2) $[\text{MnO}_4]^-$, CrO₂Cl₂
(3) $[\text{Fe}(\text{CN})_6]^{3-}$, $[\text{Co}(\text{CN})_3]$ (4) $[\text{NiCl}_4]^{2-}$, $[\text{CoCl}_4]^-$
76. Metal hydrides are ionic, covalent or molecular in nature. Among LiH, NaH, KH, RbH, CsH the correct order of increasing ionic character is
- (1) LiH > NaH > CsH > KH > RbH (2) LiH < NaH < KH < RbH < CsH
(3) RbH > CsH > NaH > KH > LiH (4) NaH > CsH > RbH > LiH > KH
77. Be and Mg have zero value of electron affinity, because
- (1) Be and Mg have [He]2s² and [Ne]3s² configuration respectively.
(2) 2s and 3s orbital are filled to their capacity.
(3) Be and Mg are unable to accept electron.
(4) All the above are correct.
78. The ground state electronic configurations of the elements, U, V, W, X and Y (these symbols do not have any chemical significance) are as follows
- U - 1s²2s²2p³
V - 1s²2s²2p⁶3s¹
W - 1s²2s²2p⁶3s²3p²
X - 1s²2s²2p⁶3s²3p⁶3d⁵4s²
Y - 1s²2s²2p⁶3s²3p⁶3d¹⁰4s²4p⁶
- Determine which sequence of elements satisfy the following statements
- (i) Element-has highly reactive nature
(ii) Element-is d-block element most likely to form coloured ionic compounds
(iii) Element-has largest atomic radius
(iv) Element-forms only acidic oxide
- (1) V W Y U (2) V X Y W (3) V W Y X (4) V X W U

79. The pair of amphoteric hydroxides is :-
 (1) $\text{Al}(\text{OH})_3$, LiOH (2) $\text{Be}(\text{OH})_2$, $\text{Mg}(\text{OH})_2$
 (3) $\text{B}(\text{OH})_3$, $\text{Be}(\text{OH})_2$ (4) $\text{Be}(\text{OH})_2$, $\text{Zn}(\text{OH})_2$
80. Interhalogen compounds are more reactive than the individual halogen because:
 (1) Two halogens are present in same group
 (2) They are more ionic
 (3) Their bond energy is less than the bond energy of the halogen molecule except F_2
 (4) They carry more energy
81. $\text{BeCl}_2 + \text{LiAlH}_4 \longrightarrow \text{X} + \text{LiCl} + \text{AlCl}_3$
 (1) X is lithium hydride (2) X is BeH_2
 (3) X is $\text{BeCl}_2 \cdot 2\text{H}_2\text{O}$ (4) X is LiH
82. Unlike PbCl_4 , PbI_4 and PbBr_4 are not found because
 (1) bromine and iodine are more electronegative than chlorine.
 (2) Iodine and bromine are smaller in size.
 (3) Larger iodine and bromine are able to reduce Pb^{4+} to Pb^{2+} or Pb .
 (4) The statement is incorrect.
83. Which does not give oxygen on heating?
 (A) HgO (2) KMnO_4 (3) KClO_3 (4) $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$
84. Which of the following compounds is not yellow coloured?
 (1) $\text{Zn}_2(\text{Fe}(\text{CN})_6)$ (2) $\text{K}_3[\text{Co}(\text{NO}_2)_6]$
 (3) $(\text{NH}_4)_3[\text{As}(\text{Mo}_3\text{O}_{10})_4]$ (4) BaCrO_4
85. Which of the following compounds is metallic and ferromagnetic?
 (1) CrO_2 (2) VO_2 (3) MnO_2 (4) TiO_2
86. On heating $\text{K}_2\text{Cr}_2\text{O}_7$ with NaCl and conc. H_2SO_4 , the gas liberated is CrO_2Cl_2 . The fumes when passed through NaOH solution form
 (1) CrO_4^{2-} (2) CrCl_3 (3) Cr_2O_3 (4) CrOCl_2
87. Which of the following has the maximum number of unpaired d – electrons?
 (1) Fe^{2+} (2) Cu^+ (3) Zn (4) Ni^{3+}
88. Among the following pair of ions, the lower oxidation state in aqueous solution is more stable in
 (1) V^{2+} , VO^{2+} (2) Cr^{2+} , Cr^{3+} (3) Ti^+ , Ti^{3+} (4) Cu^+ , Cu^{2+}
89. CuSO_4 when reacts with KCN forms CuCN which is insoluble in water. It is soluble in excess of KCN , due to formation of which of the following complex
 (1) $\text{K}_2[\text{Cu}(\text{CN})_4]$ (2) $\text{K}_3[\text{Cu}(\text{CN})_4]$
 (3) CuCN_2 (4) $\text{Cu}[\text{KCu}(\text{CN})_4]$
90. Which of the following is magnetite?
 (1) Fe_2CO_3 (2) Fe_2O_3 (3) Fe_3O_4 (4) $\text{Fe}_2\text{O}_3 \cdot 3\text{H}_2\text{O}$

PACE-IIT & MEDICAL

ANDHERI / BORIVALI / DADAR / CHEMBUR / THANE / MULUND / NERUL / POWAI

IIT – JEE - 2019

TW TEST (3 Yrs.)

TOPIC: FULL IN - ORGANIC

DATE:17/10/18

ANSWER KEY

1.	(3)	2.	(3)	3.	(2)	4.	(1)	5.	(3)	6.	(4)	7.	(4)
8.	(3)	9.	(4)	10.	(3)	11.	(3)	12.	(2)	13.	(3)	14.	(4)
15.	(4)	16.	(2)	17.	(2)	18.	(4)	19.	(2)	20.	(1)	21.	(1)
22.	(1)	23.	(4)	24.	(3)	25.	(1)	26.	(2)	27.	(2)	28.	(1)
29.	(3)	30.	(3)	31.	(2)	32.	(1)	33.	(2)	34.	(3)	35.	(3)
36.	(4)	37.	(3)	38.	(3)	39.	(1)	40.	(3)	41.	(4)	42.	(1)
43.	(2)	44.	(4)	45.	(2)	46.	(2)	47.	(4)	48.	(4)	49.	(1)
50.	(1)	51.	(3)	52.	(4)	53.	(4)	54.	(2)	55.	(4)	56.	(4)
57.	(4)	58.	(2)	59.	(3)	60.	(4)	61.	(4)	62.	(1)	63.	(2,4)
64.	(3)	65.	(1)	66.	(3)	67.	(1)	68.	(2)	69.	(4)	70.	(2)
71.	(2)	72.	(3)	73.	(1)	74.	(3)	75.	(2)	76.	(2)	77.	(4)
78.	(2)	79.	(4)	80.	(3)	81.	(2)	82.	(3)	83.	(4)	84.	(1)
85.	(1)	86.	(1)	87.	(1)	88.	(3)	89.	(2)	90.	(3)		